

**Common Polyatomic Ions**

Name(s)	Formula	Name(s)	Formula
ammonium	NH <sub>4</sub> <sup>+</sup>	iodate	IO <sub>3</sub> <sup>-</sup>
acetate	CH <sub>3</sub> COO <sup>-</sup> C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> <sup>-</sup>	nitrate	NO <sub>3</sub> <sup>-</sup>
bromate	BrO <sub>3</sub> <sup>-</sup>	nitrite	NO <sub>2</sub> <sup>-</sup>
carbonate	CO <sub>3</sub> <sup>2-</sup>	oxalate	C <sub>2</sub> O <sub>4</sub> <sup>2-</sup>
chlorate	ClO <sub>3</sub> <sup>-</sup>	perchlorate	ClO <sub>4</sub> <sup>-</sup>
chlorite	ClO <sub>2</sub> <sup>-</sup>	periodate	IO <sub>4</sub> <sup>-</sup>
chromate	CrO <sub>4</sub> <sup>2-</sup>	permanganate	MnO <sub>4</sub> <sup>-</sup>
cyanide	CN <sup>-</sup>	peroxide	O <sub>2</sub> <sup>2-</sup>
dichromate	Cr <sub>2</sub> O <sub>7</sub> <sup>2-</sup>	phosphate	PO <sub>4</sub> <sup>3-</sup>
hydrogen carbonate bicarbonate	HCO <sub>3</sub> <sup>-</sup>	phosphite	PO <sub>3</sub> <sup>3-</sup>
hydrogen sulfate bisulfate	HSO <sub>4</sub> <sup>-</sup>	silicate	SiO <sub>4</sub> <sup>4-</sup>
hydrogen phosphate biphosphate	HPO <sub>4</sub> <sup>2-</sup>	sulfate	SO <sub>4</sub> <sup>2-</sup>
hydroxide	OH <sup>-</sup>	sulfite	SO <sub>3</sub> <sup>2-</sup>
hypochlorite	ClO <sup>-</sup>	thiocyanate	SCN <sup>-</sup>
		thiosulfate	S <sub>2</sub> O <sub>3</sub> <sup>2-</sup>

**Other Ions**

copper (I)	cuprous	Cu <sup>+</sup>
copper (II)	cupric	Cu <sup>2+</sup>
iron (II)	ferrous	Fe <sup>2+</sup>
iron (III)	ferric	Fe <sup>3+</sup>
lead (II)	plumbous	Pb <sup>2+</sup>
lead (IV)	plumbic	Pb <sup>4+</sup>
mercury (I)	mercurous	Hg <sub>2</sub> <sup>2+</sup>
mercury (II)	mercuric	Hg <sup>2+</sup>
tin (II)	stannous	Sn <sup>2+</sup>
tin (IV)	stannic	Sn <sup>4+</sup>